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principle, very simply stated, is: start at the lowest energy level and build up to the higher energy levels only after the lowest ...

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Now consider

hydrogen

chloride (HCl).

Hydrogen has an electronegativity of 2.1, and

chlorine has an

electronegativity of 3.0. The

electron pair

that is bonding

HCl together

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Key

According to the modern atomic model, electrons orbit the atom's nucleus at specific levels, or shells.

Electrons in the outermost shell, which are involved in

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bonding, are known as valence electrons. Since atoms are more stable when their outer shell is filled, they tend to lose, gain, or share electrons to complete their outer shell.

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In atomic physics and quantum chemistry, the electron configuration is the distribution of electrons of an atom or molecule in

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atomic or

molecular

orbitals. For

example, the

electron

configuration of

the neon atom is

$1s^2 2s^2 2p^6$ ,

using the

notation

explained below.

Electronic

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describe each

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