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Brønsted-Lowry
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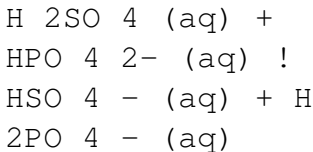
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Identify each species in the following equation as either the Brønsted-Lowry acid, the Brønsted-Lowry base, the conjugate acid, or the conjugate base. Identify the conjugate acid-base pairs

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in the reaction.



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Strong Acids:

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Answer in

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What is an acid?

2. What is the

symbol for a

hydronium ion?

3. Copy the

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Acids And

chart on pg. 241

4. What are some general properties of acids?
5. What is an indicator?
6. What is one the most common indicators used?
- 7.

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There are several sets of definitions used to distinguish between acids and bases, but Boyle's method is easy and practical. Acids tend to be sour-tasting (don't taste them!), they are

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corrosive, and they turn litmus paper red. Bases are slippery-feeling, soapy-tasting, and turn litmus paper blue.

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In fact there are only two amino acids which have only

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one sequence of bases to code for them - methionine (Met) and tryptophan (Trp). You have probably noticed that three codons don't have an amino acid written beside them, but say "stop" instead. For

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obvious reasons
these are known
as stop codons.
We'll leave
talking about
those until we
have looked at
the way the code
works in
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Key is an

example of a

strong acid.

Sodium hydroxide

(NaOH) is an

example of a

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Strong base. 8.

According to the definitions of acids and bases devised by

Lewis: acids are electron pair acceptors and bases are electron pair donors. acids are electron pair donors and

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Understand the terms

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diprotic,

triprotic, and

polyprotic.

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